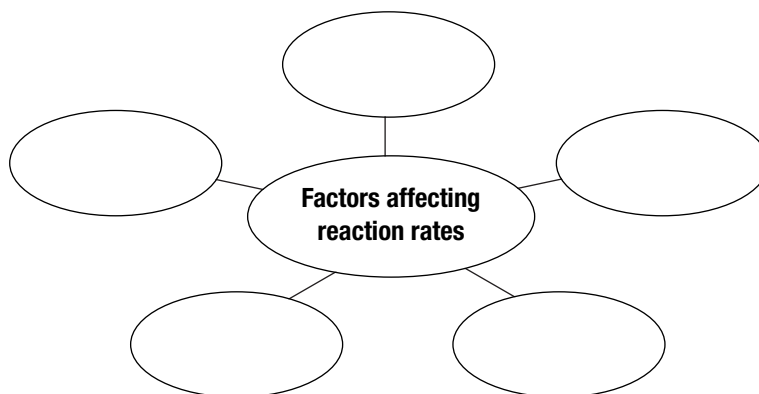


Chapter 7 Chemical Reactions

Section 7.4 Reaction Rates**(pages 212–215)***This section discusses the factors that affect reaction rates.***Reading Strategy (page 212)**

Building Vocabulary As you read, complete the web diagram below with key terms from this section. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

**Reactions Over Time (page 212)**

- Any change that happens over time can be expressed as a(n) _____.
- What is a reaction rate? _____

Factors Affecting Reaction Rates (pages 213–215)

- Is the following sentence true or false? One way to observe the rate of a reaction is to observe how fast products are being formed.

- Is the following sentence true or false? The rate of any reaction is a constant that does not change when the reaction conditions change.

- Generally, an increase in temperature will _____ the reaction rate.
- Is the following sentence true or false? Storing milk in a refrigerator stops the reactions that would cause the milk to spoil.

- How does an increase in surface area affect the exposure of reactants to one another? _____

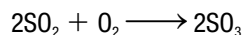
Chapter 7 Chemical Reactions

8. Why does increasing the surface area of a reactant tend to increase the reaction rate? _____

9. Stirring the reactants in a reaction mixture will generally _____ the reaction rate.
10. Is the following sentence true or false? Increasing the concentration of the reactants will generally slow down a chemical reaction. _____
11. Is the following sentence true or false? A piece of material dipped in a concentrated dye solution will change color more quickly than in a dilute dye solution. _____
12. Why does an increase in pressure speed up the rate of a reaction involving gases? _____

13. What is a catalyst? _____

14. Circle the letters of the sentences that correctly identify why chemists use catalysts.
- a. to speed up a reaction
 - b. to enable a reaction to occur at a higher temperature
 - c. to slow down a reaction
 - d. to enable a reaction to occur at a lower temperature
15. Is the following sentence true or false? Because a catalyst is quickly consumed in a reaction, it must be added to the reaction mixture over and over again to keep the reaction going. _____
16. Identify where the catalyst V_2O_5 should go in the formula shown and write it in the correct location.



17. Circle the letter of the correct answer. In the reaction represented by the equation $2H_2O_2 \xrightarrow{Pt} 2H_2O + O_2$, which substance acts as a catalyst?
- a. H_2O_2
 - b. Pt
 - c. H_2O
 - d. O_2
18. One way that a catalyst can lower the energy barrier of a reaction is by providing a surface on which the _____ can come together.